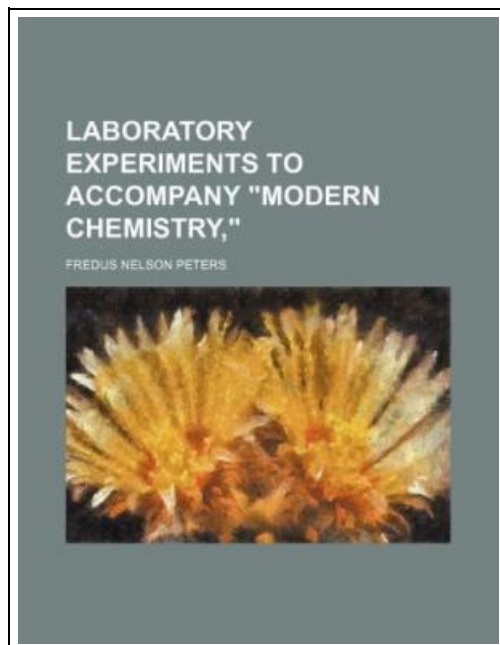


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1902 Excerpt: .4 g. sodium hydroxide. 8. Measure out accurately into a beaker 25 cc. of the alkali solution to be tested, and add one drop of phenol phthalein or some other indicator. From a burette run in the solution of oxalic acid slowly till neutralization is effected. Note the number of cubic centimeters used. From this we can determine the amount of alkali. By the equation above, we see that 6.3 g. of acid combine with 4 g. of the alkali. But 6.3 g. of acid are contained in 1000 cc. of the solution. Then--1000 cc. acid: 4 g. NaOH:: m (no cc. acid used): x. But x is the amount of sodium hydroxide in 25 cc.; from this find the amount in 11., 1000 cc. For other problems of similar character, see Modern Chemistry, page 169, or any work in volumetric analysis. CHAPTER XI SULPHUR AND SOME COMPOUNDS 1. Various Forms of Sulphur Crystals.--Add a little powdered brimstone to 1 or 2 cc. of carbon disulphide in a test-tube, and shake. When the sulphur has dissolved, pour out the solution upon a watch glass, and set it to crystallize in as cool a place as possible. Later, examine the crystals with a lens, and note their form. Make a drawing of one of the crystals. 2. If this experiment is not done in the laboratory, write up from work in the class room. Fill a small sand crucible nearly full of sulphur, and heat till the...



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